Additions and Corrections

Nucleophilic Displacement on Sulfur. The Inversion of Sulfoxide Configuration [*J. Am. Chem. Soc.*, **87**, 5404 (1965)]. By CARL R. JOHNSON and DAVID MCCANTS, JR. Department of Chemistry, Wayne State University, Detroit, Michigan.

Professor Kurt Mislow has pointed out a dramatic change in rotation of benzyl *p*-tolyl sulfoxide (I) with change in solvent. A sample of I, found to have $[\alpha]^{20}\text{D}$ +236° (acetone) and therefore 93.7% optically pure based on Stirling's maximum value of $[\alpha]^{20}\text{D}$ +252° (acetone), was found to have $[\alpha]^{20}\text{D}$ +101° (chloroform). The rotation of optically pure I is then presumably 101°/0.937 or 107.8° (chloroform). It follows that our sulfoxide (I) is approximately 88% optically pure rather than the 37.5% as suggested on page 5407.

The Preparation and Properties of *trans*-Cinnamoyl-Papain [*J. Am. Chem. Soc.*, **88**, 5871 (1966)]. By LEWIS J. BRUBACHER and MYRON L. BENDER. Northwestern University, Evanston, Illinois.

The headings in Table IV (page 5878) are incorrect. The table should read as shown below

Nucleophile	$\frac{k_{4\text{II}}\text{ cinnamoyl-papain}}{k_{4\text{II}}\text{ cinnamoyl-}\alpha\text{-chymotrypsin}}$	
	Actual	Normalized
Water	0.30	1[c]
Ammonia	>143.[a]	>480
Methanol	0.23	1[c]
Methylamine	155.[b]	670

(3)-1,2-Dicarbollyl Complexes of Nickel(III) and Nickel(IV) [*J. Am. Chem. Soc.*, **89**, 470 (1967)]. By LESLIE F. WARREN, JR., and M. FREDERICK HAWTHORNE. Department of Chemistry, University of California, Riverside, California 92502.

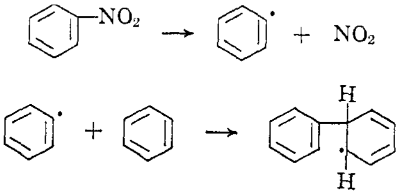
The extinction coefficients for the ultraviolet–visible spectrum of $Ni(B_9C_2H_{11})_2$ (II) should be changed to $\lambda_{max}(\epsilon)$: 293 (37,000) and 425 mμ (2200).

The Configuration of Pyrazoline Azomethine Dyes [*J. Am. Chem. Soc.*, **89**, 580 (1967)]. By P. J. S. PAUWELS. Research Laboratories, Kodak Ltd., Kirkby, Liverpool, England.

In the abstract, lines 2 and 3 should read 3-methyl and 3-substituent instead of 4-methyl and 4-substituent, respectively.

Arylation by Aromatic Nitro Compounds at High Temperatures [*J. Am. Chem. Soc.*, **89**, 724 (1967)]. By ELLIS K. FIELDS and SEYMOUR MEYERSON. Amoco Chemicals Corporation and American Oil Company, Whiting, Indiana.

The dots of the phenyl radicals were omitted from the first two lines of equations on page 725. The equations should read as shown below.



Book Reviews

Metal π-Complexes. Volume 1. Complexes with Di- and Oligo-Olefinic Ligands. By E. O. FISCHER, Professor of Inorganic Chemistry, and H. WERNER, Research Assistant, both of the Technische Hochschule, Munich, Germany. American Elsevier Publishing Co., Inc., 52 Vanderbilt Ave., New York, N. Y. 1966. 246 pp. 15.5 × 23 cm. $16.00.

This book is a translation and extension of an earlier (1963) book in German by the same authors with essentially the same title. Three new sections have been added to Chapter V of the older version and a new Chapter (VI) deals with π-allyl complexes. Apparently between editions the authors began to think in terms of a series of books with the general title, "Metal π-Complexes." The present book has now become Volume 1 in the series, and Volume 2, "Complexes with Mono-Olefinic Ligands," is stated to be in preparation. Such plans, however, have not resulted in any obvious departures from the structure or organization of the German book. The present volume contains 270 more references than the earlier one, all arranged alphabetically. Many of these are grafted onto the older list since reference 450 to Zeise is followed by 451 which begins the alphabetical listing again. It seems to this reviewer that listing references alphabetically at the end of a book has little merit unless the page in the text on which the reference appears is also included. This is a particularly poor practice when, as in the present case, there is no authors' index.

The book is organized on the basis of the structure of the ligand, and thus complexes of all metals with the same ligand are treated together. Such an organization brings up the question as to where one should expect to find a complex containing two different olefinic ligands bonded to a metal, for example, a norbornadiene and cyclopentadienyl simultaneously bonded to nickel. Some dispersal and repetition are intrinsic to such a classification, but probably any classification scheme in such a complex field would probably have drawbacks. One may question the logic of including π-allyl complexes in a volume dealing with di- and oligo-olefinic ligands, and where exactly oligo ends and poly begins are left undefined.

This book is more a reference text than one which can be "read" easily by the uninitiated. Most of the material is taken directly from the literature without an attempt to evaluate or integrate it. A substantial portion of the work covered was performed in the authors' Munich laboratory (about 50 papers by the senior author are cited), and unquestionably the authors have had personal contact with all the leaders in the field. As a result, this is a very thorough and authentic book, written with accuracy and clarity. Anyone interested in knowing whether a particular π-olefinic complex has been prepared and if so, what the properties of it are, should first consult this small, valuable, but relatively expensive book.

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